

## **AMENDMENTS TO THE SPECIFICATION**

**Replace the third paragraph beginning on page 9 with the following:**

Ratios of isotopic concentrations are preferred as they present two distinct advantages over individual concentrations. First, isotopic ratios can be more reproducibly measured than compositions. Second, isotopic ratios may not be modified by non-nuclear physical or chemical processes or explosions such that ratios will remain intact through subsequent chemical reactions, including ~~tire wear and explosions~~ fires and explosions.”

**Replace the first full paragraph on page 10 with the following:**

Further, isotopic concentrations provide stable isotopic identifications which are highly specific. Elements which have more than one stable isotope are numerous. Of the 83 known non-radioactive elements known to exist on earth, 62 have more than one stable isotope, and 40 have more than two stable isotopes. The element tin (Sn) has the largest number of stable isotopes for any single element. Among the 40 elements having more than two stable isotopes, there are a total of 224 stable isotopes. Although a few of the 224 ~~252~~ stable isotopes are slightly radioactive, they have very long lives and are present in many naturally occurring elements. Thus, as will be seen, the stable isotopic identifications of the invention are numerous and provide a ready and available means by which any product (including all pharmaceutical phases APIs, drug products, excipients of drug products and/or impurities of drug products) may be readily identified.

**Replace the first full paragraph on page 13 with the following:**

Additionally, the “error of identification” can be reduced or, the “precision of identification” can be increased by choosing more than one isotopic concentration. There are a

total of 13 if one limits the stable isotopic identification of the invention to the common light elements. Reduced error can be accomplished by using any number of the total of 252 available stable isotopes of elements having two or more stable isotopes.